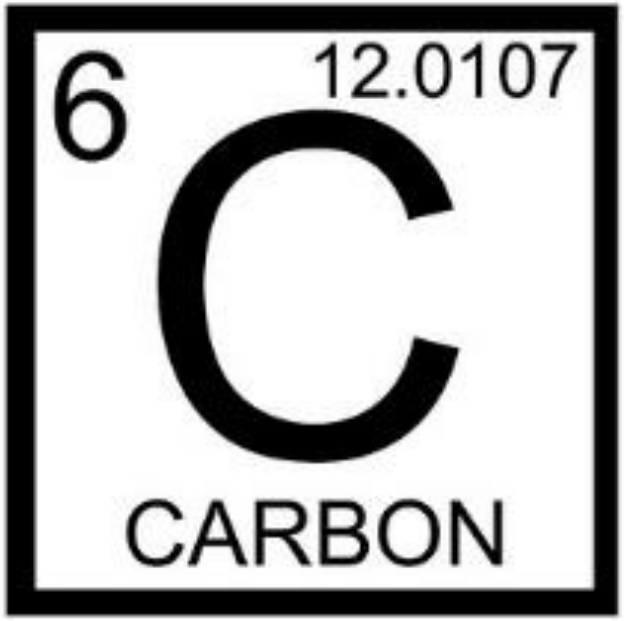
How to Draw Lewis Structures

Lewis Structures

1) Find your element on the periodic table.
2) Determine the number of valence electrons.
3) This is how many electrons you will draw.



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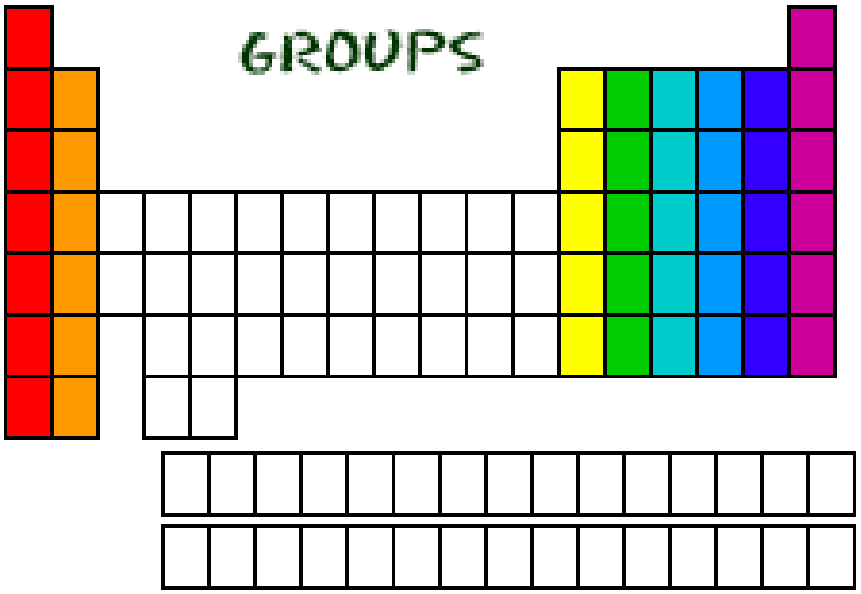

- Find out which group (column) your element is in.
- This will tell you the number of valence electrons your element has.
- You will only draw the valence electrons.

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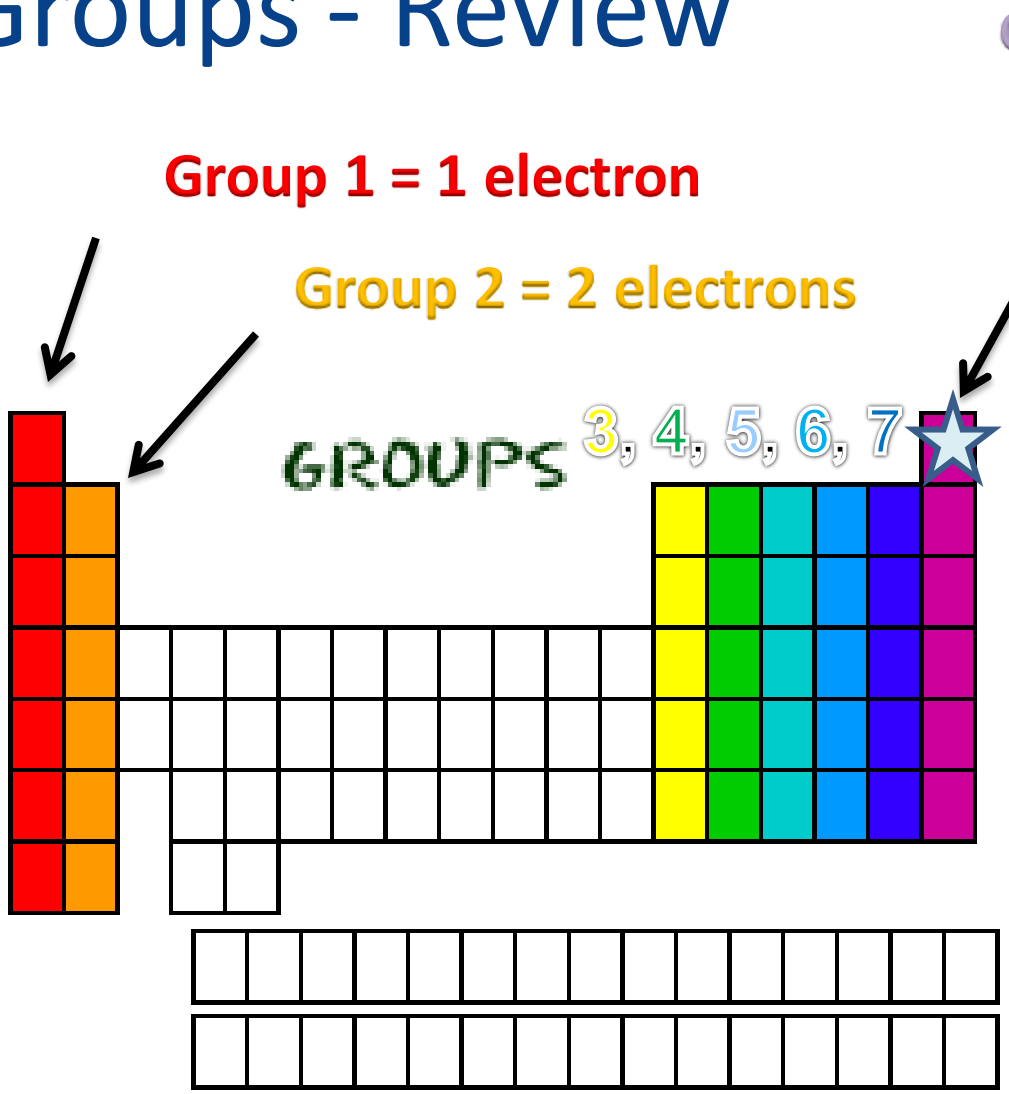
Groups - Review



- **Each column is called a "group"**
- **Each element in a group has the same number of electrons in their outer orbital, also known as "shells".**
- **The electrons in the outer shell are called "valence electrons"**

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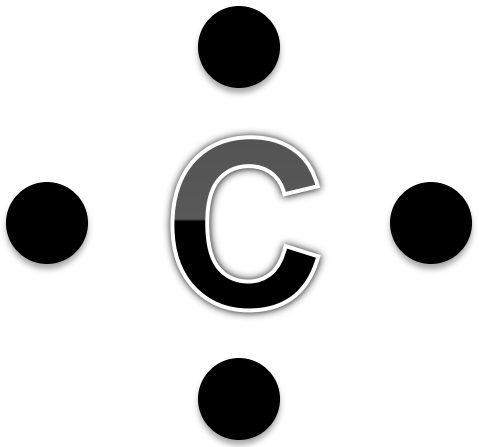
Lewis Structures



1) Write the element symbol.
2) Carbon is in the 4th group, so it has 4 valence electrons.
3) Starting at the right, draw 4 electrons, or dots, counter-clockwise around the element symbol.

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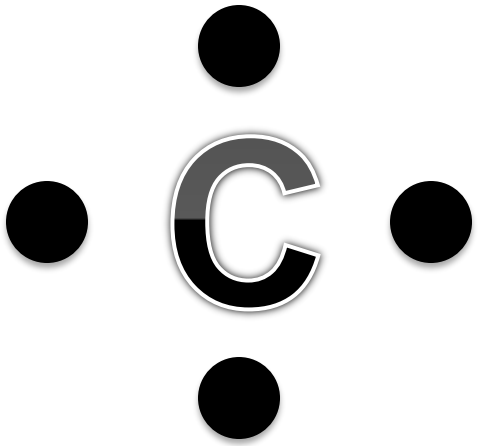

1) Check your work.
2) Using your periodic table, check that Carbon is in the 4^{th} group.
3) You should have 4 total electrons, or dots, drawn in for Carbon.

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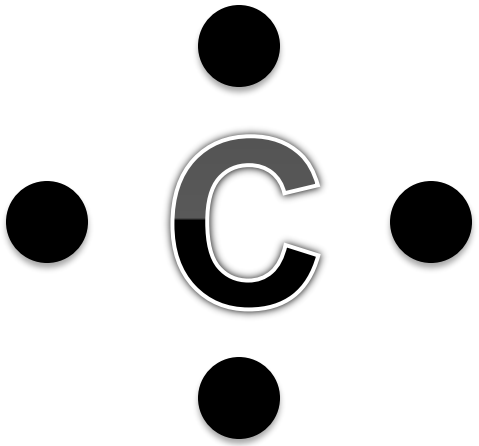

On your worksheet, try these elements on your own:

a) H
b) P
c) Ca
d) Ar
e) Cl
f) Al

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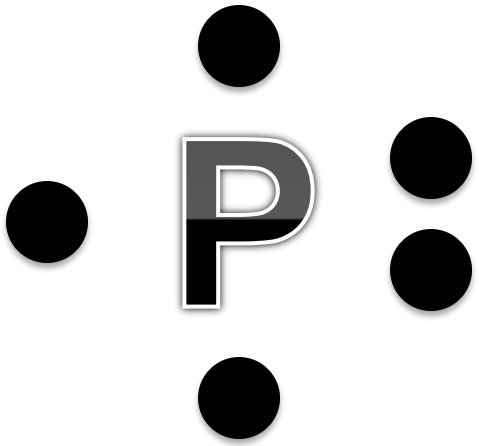

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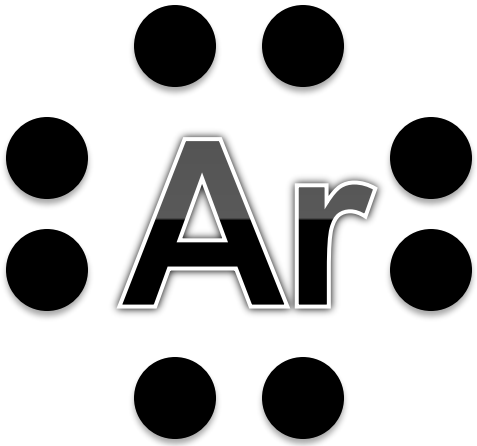

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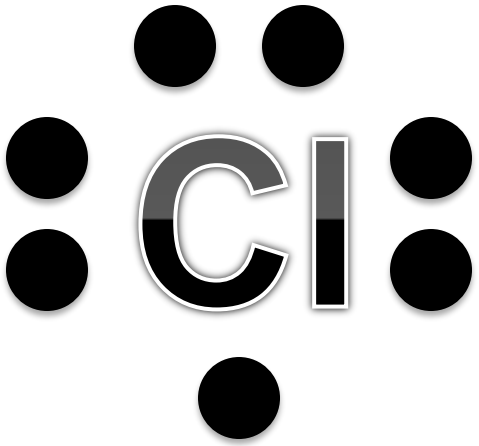

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